CH 15 Equilibrium Study Guide

Dynamic Equilibrium

Dynamic Equilibrium - forward & reverse reaction rates are equal

- ASSUME all rxns are ELEMENTAL (1 step!)
- Reversible reaction something constantly produced & consumed
 - Indicated by

No net change in reactant or product concentrations

Equilibrium Constant, Ke

For a general equation:

$$aA + bB \xrightarrow{k_t} aD + eE$$

$$K_c = \frac{k_t}{k_c} = \frac{[D]^d [E]^e}{[A]^e [B]^b}$$

A. The equilibrium-constant depends only on the stoichiometry of the reaction, not the mechanism.

B. DOES NOT depend on initial concentrations because only the RATIO of concentrations at equilibrium is the same

C. K has NO units!

D. ONLY consider gas & aqueous phases!!

-OMIT solids & liquids

Significance of K

Value of K indicates if more products or reactants are present at equilibrium

- . If K is large (>1): Products favored Equilibrium lies to the right
- . If K is small (<1): Reactants favored Equilibrium lies to the left

Equilibrium Constant, Ko

For a general equation:

Only if ALL phases are gases!

$$aA + bB \xrightarrow{k_f} dD + eE K_p = \frac{(P_D)^r(P_E)^r}{(P_A)^r(P_B)^r}$$

Relating Ko and Ko

$$K_{p} = K_{c}(RT)^{\Delta n}$$
; $\Delta n = [\Sigma \text{ (moles GAS in product)} - \Sigma \text{ (moles GAS in reactant)}]$

Dynamic Equilibrium Study Guide

Bruce Wingerd, Patty Bostwick Taylor

Dynamic Equilibrium Study Guide:

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Table of Contents Dynamic Equilibrium Study Guide

- 1. Understanding the eBook Dynamic Equilibrium Study Guide
 - The Rise of Digital Reading Dynamic Equilibrium Study Guide
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Dynamic Equilibrium Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Dynamic Equilibrium Study Guide
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Dynamic Equilibrium Study Guide
 - Personalized Recommendations
 - Dynamic Equilibrium Study Guide User Reviews and Ratings

- Dynamic Equilibrium Study Guide and Bestseller Lists
- 5. Accessing Dynamic Equilibrium Study Guide Free and Paid eBooks
 - o Dynamic Equilibrium Study Guide Public Domain eBooks
 - Dynamic Equilibrium Study Guide eBook Subscription Services
 - Dynamic Equilibrium Study Guide Budget-Friendly Options
- 6. Navigating Dynamic Equilibrium Study Guide eBook Formats
 - o ePub, PDF, MOBI, and More
 - Dynamic Equilibrium Study Guide Compatibility with Devices
 - Dynamic Equilibrium Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Dynamic Equilibrium Study Guide
 - Highlighting and Note-Taking Dynamic Equilibrium Study Guide
 - Interactive Elements Dynamic Equilibrium Study Guide
- 8. Staying Engaged with Dynamic Equilibrium Study Guide
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Dynamic Equilibrium Study Guide
- 9. Balancing eBooks and Physical Books Dynamic Equilibrium Study Guide
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Dynamic Equilibrium Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Dynamic Equilibrium Study Guide
 - Setting Reading Goals Dynamic Equilibrium Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Dynamic Equilibrium Study Guide
 - Fact-Checking eBook Content of Dynamic Equilibrium Study Guide
 - $\circ \ Distinguishing \ Credible \ Sources$

- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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