



## Unit 6 Test Study Guide

January 13th, 2022

### Brensted-Lowry Acid/Base Theory

- Acids: Proton ( $H^+$ ) donor
- Bases: Proton ( $H^+$ ) acceptor
- Conjugate Acids-Base Pair: Two substances that differ only by the presence of an  $H^+$ 
  - Conjugate Acid: Formed when a proton is added to a base (positive charge)
    - Add a Hydrogen to the formula and add +1 charge
  - Conjugate Base: Formed when a proton is removed from an acid (negative charge)
    - Take away a Hydrogen from the formula and subtract -1 charge
- Amphoteric: The ability for a compound to act as an acid OR as a base (EX:  $H_2O$ )
- Acids are classified based on how many  $H^+$  are present
  - Monoprotic: 1  $H^+$
  - Diprotic: 2  $H^+$
  - Triprotic: 3  $H^+$
  - Diprotic and triprotic acids are also amphoteric because they have more than one proton that they can lose

### Properties of Strong and Weak Acids/Bases

Strong Acids/Bases	Weak Acids/Bases
<ul style="list-style-type: none"><li>• Strong electrolytes</li><li>• Ionize completely in water</li><li>• Equilibrium is not present<ul style="list-style-type: none"><li>◦ Very unlikely that the ions reform into products</li></ul></li><li>• Very weak attraction between ions<ul style="list-style-type: none"><li>◦ Split apart easily</li></ul></li></ul>	<ul style="list-style-type: none"><li>• Weak electrolytes</li><li>• Does not ionize completely in water</li><li>• Equilibrium exists</li><li>• Strong attraction between ions<ul style="list-style-type: none"><li>◦ Do not split apart easily</li></ul></li></ul>
<p style="text-align: center;"><b>Strong acid</b></p> 	<p style="text-align: center;"><b>Weak acid</b></p> 

# Equilibrium And Acid Base Study Guide

**David Margolese, David W. Oxtoby**



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