Name:	
Chemistry Regents	

Date:\_\_\_\_\_ Leahy

#### Fifty Frequently Forgotten Facts of Chemistry

#### I. ATOMIC STRUCTURE & NUCLEAR CHEMISTRY

 Protons are +1 each with a mass of 1 amu each, the number of protons = atomic number, nuclear charge = + (= protons). [Periodic Table]

a) How many protons are there in a nucleus of Kr-85 ?

b) What is the nuclear charge of an atom of Br?\_\_\_\_\_

c) What is the mass of the protons in a nucleus of O-15?

2) Neutrons are neutral with a mass of 1 amu each,  $\pi$  neutrons = mass  $\pi$  - atomic number. Isotopes = atoms of the same element (same atomic  $\pi$ ) but different  $\pi$  of neutrons (mass  $\pi$ ). [Periodic Table]

a) How many neutrons are there in the nucleus of 10 Fe?

b) Circle the two nuclei that are isotopes of each other: 11,0 15,N 15,0 15,F

3) Electrons are each -1 with a mass that is VERY, VERY tiny compared to the mass of a proton or neutron.

a) Which particle has a mass that is 1/1836" the mass of a proton?

1) "He 2) "H 3) "se

4) <sup>7</sup>am

4) Natural Decay: Parent Nuclide -> Decay particle + daughter nuclide [Tables N and O]

a) Write the decay for U-238:

b) Write the decay for K-37:

c) Write the decay for P-32:

5) <u>Artificial Transmutation</u> is when a relatively stable nucleus is impacted by a particle bullet at high speeds and becomes an unstable nucleus of a different element. <u>Nuclear fission</u> occurs when nuclei of U-235 or Pu-239 are impacted by a neutron and split into two smaller nuclei and more neutrons. <u>Nuclear fusion</u> occurs when two small nuclei of hydrogen combine at high temperatures and pressures to form larger nuclei of helium. Both fission and fusion convert mass into a huge amount of energy.

Given the nuclear reactions:

a) Which reaction represents natural decay?\_\_\_\_

b) Which reaction represents artificial transmutation?\_\_\_\_\_

c) Which reaction represents nuclear fission?\_\_\_\_

d) Which reaction represents nuclear fusion?\_\_\_\_\_

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M Mark

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