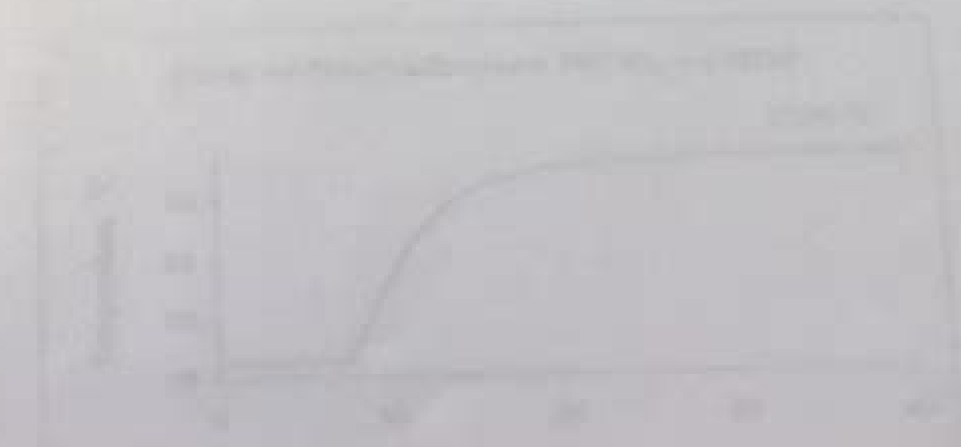


4. In accordance with Hess's Law, manipulate the following equations and their ΔH_{rxn} 's to calculate ΔH_{rxn} for the reaction $\text{HClO}_4 + \text{NH}_3 \rightarrow \text{NH}_4\text{ClO}_4$.



_____	$\Delta H_{\text{rxn}} = \text{_____ kJ/mol}$
_____	$\Delta H_{\text{rxn}} = \text{_____ kJ/mol}$
_____	$\Delta H_{\text{rxn}} = \text{_____ kJ/mol}$
$\text{HClO}_{4(aq)} + \text{NH}_3_{(aq)} \rightarrow \text{NH}_4\text{ClO}_{4(aq)}$	$\Delta H_{\text{rxn}} = \text{_____ kJ/mol}$



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