

## Experiment 24

### Advance Study Assignment: Molar Mass of an Acid

1. 7.2 mL of 6.0 M NaOH are diluted with water to a volume of 400.0 mL. You are asked to find the molarity of the resulting solution.

- a. First find out how many moles of NaOH there are in 7.2 mL of 6.0 M NaOH. Use Equation 1. Note that the volume must be in liters.

\_\_\_\_\_ moles

- b. Since the total number of moles of NaOH is not changed on dilution, the molarity after dilution can also be found by Equation 1, using the final volume of the solution. Calculate that molarity.

\_\_\_\_\_ M

2. In an acid-base titration, 21.16 mL of an NaOH solution are needed to neutralize 20.04 mL of a 0.0997 M HCl solution. To find the molarity of the NaOH solution, we can use the following procedure:

- a. First note the value of  $M_{H^+}$  in the HCl solution.

\_\_\_\_\_ M

- b. Find  $M_{OH^-}$  in the NaOH solution. (Use Eq. 3.)

\_\_\_\_\_ M

- c. Obtain  $M_{NaOH}$  from  $M_{OH^-}$ .

\_\_\_\_\_ M

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# Experiment 24 Advance Study Assignment Molar Mass Of An Acid Answers

**BM King**



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