



## Limiting Reactant

Experiment

4

A *limiting reactant* is the reagent that is completely consumed during a chemical reaction. Once this reagent is consumed the reaction stops. An *excess reagent* is the reactant that is left over once the limiting reagent is consumed. The maximum theoretical yield of a chemical reaction is dependent upon the limiting reagent thus the one that produces the least amount of product is the limiting reagent. For example, if a 2.00 g sample of ammonia is mixed with 4.00 g of oxygen in the following reaction, use stoichiometry to determine the limiting reagent.



$$2.00 \text{ g NH}_3 \times \frac{\text{mol}}{17.0 \text{ g}} \times \frac{4 \text{ mol NO}}{4 \text{ mol NH}_3} \times \frac{30.0 \text{ g}}{\text{mol}} = 3.53 \text{ g NO}$$

$$4.00 \text{ g O}_2 \times \frac{\text{mol}}{32.0 \text{ g}} \times \frac{4 \text{ mol NO}}{5 \text{ mol O}_2} \times \frac{30.0 \text{ g}}{\text{mol}} = 3.00 \text{ g NO}$$

Since the 4.00 g of  $\text{O}_2$  produced the least amount of product,  $\text{O}_2$  is the limiting reagent.

In this experiment you will be given a mixture of two ionic solids,  $\text{AgNO}_3$  and  $\text{K}_2\text{CrO}_4$ , that are both soluble in water. When the mixture is dissolved into water they will react to form an insoluble compound,  $\text{Ag}_2\text{CrO}_4$ , which can be collected and weighed. The reaction is the following:



*Precipitate* is the term used for an insoluble solid which is produced by mixing two solutions. Often the solid particles are so fine that they will pass right through the pores of a filter. Heating the solution for a while causes the particles to clump together, a process called *digesting*. The precipitate coagulates upon cooling and eventually settles to the bottom of a solution container. The clear liquid above the precipitate is called the *supernatant*. When the mixture is poured into a filter, solid is trapped by the filter and the clear liquid passes through as the *filtrate*. You will then test the filtrate to find out the limiting reagent and then weigh the precipitate collected to determine the mass of each reagent originally present in the solid mixture.

### Equipment and Reagents

|                       |                     |                 |                                |
|-----------------------|---------------------|-----------------|--------------------------------|
| Unknown solid mixture | hot plate or burner | 2 watch glasses | 0.2 M $\text{K}_2\text{CrO}_4$ |
| 2 Shell vials         | beaker tongs        | 150 mL beaker   | 0.1 M $\text{AgNO}_3$          |
| Analytical balance    | filter paper        | wash bottle     |                                |

# Experiment 7 The Limiting Reactant Gonzaga A

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