

Experiment 34 Prelaboratory Assignment

An Equilibrium Constant

Date _____ Lab Sec. _____ Name _____ Desk No. _____

1. Three parameters affect the absorbance of a sample. Which one is the focus of this experiment?

(1) Concentration of the solution

(2) Tube length

(3) Wavelength

Concentration is being focused on

2. Experimental Procedure, Part A.1, Table 34.1. A 3.00-ml. aliquot of 0.001 M NaSCN is diluted to 25.0 ml. with 0.2 M $\text{Fe}(\text{NO}_3)_3$ and 0.1 M HNO_3 .

a. How many moles of SCN^- are present?

$$\text{no. of moles of } \text{SCN}^- = \frac{0.001 \text{ M} \times 3 \text{ ml}}{1000 \text{ ml}} = 3 \times 10^{-6} \text{ mol}$$

- b. If all of the SCN^- is complexed with Fe^{3+} to form FeNCS^{2+} , what is the molar concentration of FeNCS^{2+} ?

$$3 \times 10^{-6} \text{ mol of } \text{SCN}^- \text{ produce } 3 \times 10^{-6} \text{ mol of } \text{FeNCS}^{2+}$$

Total volume of diluted solution = 25 ml = 0.025 L

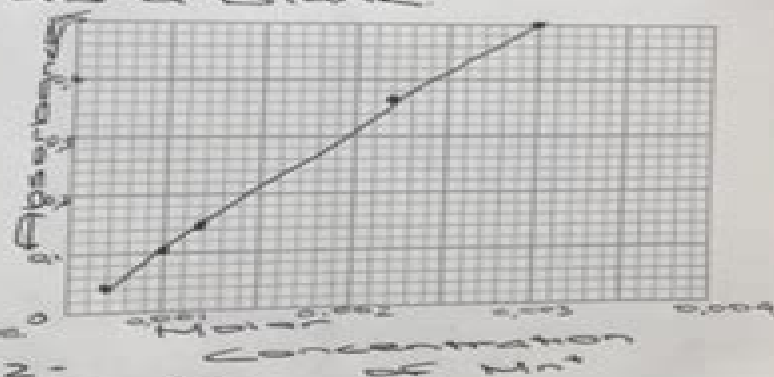
3. Experimental Procedure, Part A.1. For preparing a set of standard solutions of FeNCS^{2+} , the equilibrium molar concentration of FeNCS^{2+} is assumed to equal the initial molar concentration of the SCN^- in the reaction mixture. Why is this assumption valid? The SCN^- is the limiting reagent and thus when all of SCN^- is added its concentration no more complex formed.

4. Experimental Procedure, Part A.3. The blank solution used to calibrate the spectrophotometer is 10.0 ml. of 0.2 M $\text{Fe}(\text{NO}_3)_3$ diluted to 25.0 ml. with 0.1 M HNO_3 . Why is this solution preferred to simply using de-ionized water for the calibration? The blank prepared above also quenchess the color which must be subtracted to the correct value for the complex.

When the actual reaction is done, so, blank water cannot be used as a blank.

5. Plot the following data as absorbance versus $[\text{M}^{3+}]$ as a calibration curve:

Absorbance, A	Molar Concentration of M^{3+}
0.045	$3.0 \times 10^{-4} \text{ mol/L}$
0.097	$6.2 \times 10^{-4} \text{ mol/L}$
0.14	$9.0 \times 10^{-4} \text{ mol/L}$
0.35	$2.2 \times 10^{-3} \text{ mol/L}$
0.51	$3.2 \times 10^{-3} \text{ mol/L}$



A test solution showed a percent transmittance (%T) reading of 38.4%T. Interpret the calibration curve to determine the molar concentration of M^{3+} in the test solution.

$$\begin{aligned} \text{Absorbance} &= 2 - 10\log\%T = 2 - 10\log(38.4) = 0.416 \\ \text{Concentration of unknown solution} &= 0.416 \times 3.2 \times 10^{-3} / 0.51 = 2.61 \times 10^{-3} \text{ mol/L} \end{aligned}$$

Experiment 34 An Equilibrium Constant Answers

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1. Understanding the eBook Experiment 34 An Equilibrium Constant Answers
 - The Rise of Digital Reading Experiment 34 An Equilibrium Constant Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Experiment 34 An Equilibrium Constant Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Experiment 34 An Equilibrium Constant Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Experiment 34 An Equilibrium Constant Answers
 - Personalized Recommendations
 - Experiment 34 An Equilibrium Constant Answers User Reviews and Ratings
 - Experiment 34 An Equilibrium Constant Answers and Bestseller Lists
5. Accessing Experiment 34 An Equilibrium Constant Answers Free and Paid eBooks
 - Experiment 34 An Equilibrium Constant Answers Public Domain eBooks
 - Experiment 34 An Equilibrium Constant Answers eBook Subscription Services
 - Experiment 34 An Equilibrium Constant Answers Budget-Friendly Options
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 - Experiment 34 An Equilibrium Constant Answers Compatibility with Devices
 - Experiment 34 An Equilibrium Constant Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Experiment 34 An Equilibrium Constant Answers
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 - Interactive Elements Experiment 34 An Equilibrium Constant Answers
 8. Staying Engaged with Experiment 34 An Equilibrium Constant Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Experiment 34 An Equilibrium Constant Answers
 9. Balancing eBooks and Physical Books Experiment 34 An Equilibrium Constant Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Experiment 34 An Equilibrium Constant Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Experiment 34 An Equilibrium Constant Answers
 - Setting Reading Goals Experiment 34 An Equilibrium Constant Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Experiment 34 An Equilibrium Constant Answers
 - Fact-Checking eBook Content of Experiment 34 An Equilibrium Constant Answers
 - Distinguishing Credible Sources
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 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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