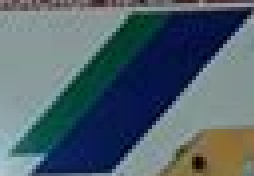


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$$\begin{aligned} &1 \text{ n}_2 + 3 \text{ f}_2 \rightarrow 2 \text{ nf}_3 \\ &2 \text{ c}_6\text{h}_{10} + 17 \text{ o}_2 \rightarrow 12 \text{ co}_2 + 10 \text{ h}_2\text{o} \\ &3 \text{ hbr} + 1 \text{ khco}_3 + 1 \text{ h}_2\text{o} + 1 \text{ kbr} + 1 \text{ co}_2 + 2 \text{ gabr}_3 + 3 \text{ na}_2\text{so}_3 + 1 \text{ ga}_2\text{so}_3 + 3 \text{ h}_2\text{o} + 3 \text{ n}_2\text{o}_3 \\ &4 \text{ naso}_3 + 3 \text{ gaz} + 3 \text{ so}_3 + 5 \text{ nf}_3 + 3 \text{ snf}_2 + 2 \text{ n}_2\text{o}_3 \end{aligned}$$
 using the equation from problem 2 above answer the following

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$$\begin{aligned} &1 \text{ n}_2 + 3 \text{ f}_2 \rightarrow 2 \text{ nf}_3 \\ &2 \text{ c}_6\text{h}_{10} + 17 \text{ o}_2 \rightarrow 12 \text{ co}_2 + 10 \text{ h}_2\text{o} \\ &3 \text{ hbr} + 1 \text{ khco}_3 + 1 \text{ h}_2\text{o} + 1 \text{ kbr} + 1 \text{ co}_2 + 2 \text{ gabr}_3 + 3 \text{ na}_2\text{so}_3 + 1 \text{ ga}_2\text{so}_3 + 3 \text{ h}_2\text{o} + 3 \text{ n}_2\text{o}_3 \\ &2 \text{ naoh} + \text{h}_2\text{so}_4 \rightarrow 2 \text{ h}_2\text{o} + \text{na}_2\text{so}_4 \end{aligned}$$
 using the following equation

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$$\text{ce} + \text{zn} + \text{cucl}_2 \rightarrow \text{zncl}_2 + \text{cu} + \text{zn} + \text{cucl}_x$$
 2 znclx 2 cu how many moles of ce zncl2 znclx 2 will be produced from 23.0 g of ce zn zn assuming ce cucl2 cuclx 2 is available in excess

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$$\begin{aligned} &1 \text{ n}_2 + 3 \text{ f}_2 \rightarrow 2 \text{ nf}_3 \\ &2 \text{ c}_6\text{h}_{10} + 17 \text{ o}_2 \rightarrow 12 \text{ co}_2 + 10 \text{ h}_2\text{o} \\ &3 \text{ hbr} + 1 \text{ khco}_3 + 1 \text{ h}_2\text{o} + 1 \text{ kbr} + 1 \text{ co}_2 + 2 \text{ gabr}_3 + 3 \text{ na}_2\text{so}_3 + 1 \text{ ga}_2\text{so}_3 + 3 \text{ h}_2\text{o} + 3 \text{ n}_2\text{o}_3 \\ &2 \text{ naoh} + \text{h}_2\text{so}_4 \rightarrow 2 \text{ h}_2\text{o} + \text{na}_2\text{so}_4 \end{aligned}$$
 using the following equation

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web so basing on my understanding of this so far an atom is an element that has the same number of protons and neutrons an ion is an element that has different numbers of protons and electrons an ion is positive when it has more protons than electrons and negative when it has more electrons than ions while isotopes are the same element with

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