

TYPES OF CHEMICAL REACTIONS

DISPLACEMENT (SINGLE AND DOUBLE)

INSTRUCTIONS:

Balance the equations below. Indicate the type of displacement reaction by writing SINGLE or DOUBLE in the space provided. Examples are provided below.



Reaction: **SINGLE** DISPLACEMENT



Reaction: **DOUBLE** DISPLACEMENT

QUESTIONS:

- 1) $__\text{Cl}_2 + __\text{KI} \rightarrow __\text{KCl} + __\text{I}_2$ _____DISPLACEMENT
- 2) $__\text{K}_2\text{CO}_3 + __\text{BaCl}_2 \rightarrow __\text{KCl} + __\text{BaCO}_3$ _____DISPLACEMENT
- 3) $__\text{Na} + __\text{MgCl}_2 \rightarrow __\text{NaCl} + __\text{Mg}$ _____DISPLACEMENT
- 4) $__\text{Al} + __\text{CuCl}_2 \rightarrow __\text{AlCl}_3 + __\text{Cu}$ _____DISPLACEMENT
- 5) $__\text{Al} + __\text{Pb(NO}_3)_2 \rightarrow __\text{Al(NO}_3)_3 + __\text{Pb}$ _____DISPLACEMENT
- 6) $__\text{AgNO}_3 + __\text{CaBr}_2 \rightarrow __\text{AgBr} + __\text{Ca(NO}_3)_2$ _____DISPLACEMENT
- 7) $__\text{Al}_2(\text{SO}_4)_3 + __\text{Ca}_3(\text{PO}_4)_2 \rightarrow __\text{AlPO}_4 + __\text{CaSO}_4$ _____DISPLACEMENT
- 8) $__\text{FeBr}_2 + __\text{K}_2\text{CO}_3 \rightarrow __\text{FeCO}_3 + __\text{KBr}$ _____DISPLACEMENT
- 9) $__\text{Cu} + __\text{Ag}_2\text{SO}_4 \rightarrow __\text{Ag} + __\text{CuSO}_4$ _____DISPLACEMENT
- 10) $__\text{CuS} + __\text{KCl} \rightarrow __\text{CuCl}_2 + __\text{K}_2\text{S}$ _____DISPLACEMENT

Double Displacement Reactions Answer Key

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