

12

THE BEHAVIOR OF GASES

Conceptual Curriculum

- Concrete concepts
- More abstract concepts or math/problem-solving

Standard Curriculum

- Core content
- Extension topics

Honors Curriculum

- ◆ Core honors content
- ◇ Options to accelerate

SECTION 12.1 THE PROPERTIES OF GASES (pages 327–328)



This section describes the properties of gas particles and explains how the kinetic energy of gas particles relates to Kelvin temperature.

► Kinetic Theory Revisited (pages 327–328)

- What theory explains the behavior of gases with respect to conditions such as temperature and pressure? kinetic theory
- If you notice that a sealed bag of potato chips bulges when placed near a sunny window, what can you hypothesize about the relationship between the temperature and pressure of an enclosed gas?
An increase in temperature causes an increase in the pressure exerted by the gas.
- List three basic assumptions of the kinetic theory about the properties of gases.
 - Gas particles are so small in relation to the distances between them that their individual volumes can be assumed to be insignificant.
 - No attractive or repulsive forces exist between gas particles.
 - Gas particles are in constant random motion, traveling in independent straight paths until they collide with each other or other objects.
- Circle the letter next to each sentence that is true concerning the compressibility of gases.
 - ☒ The large relative distances between gas particles means that there is considerable empty space between them.
 - ☒ The assumption that gas particles are far apart explains gas compressibility.
 - ☒ Compressibility is a measure of how much the volume of matter decreases under pressure.
 - ☐ Energy is released by a gas when it is compressed.
- Look at Figure 12.1 on page 327. How does an automobile air bag protect the crash dummy from being broken as a result of impact?
The gases used to inflate the airbag are compressible and are able to absorb a considerable amount of energy.

Directed Reading For Content Mastery Section 3

Behavior Of Gases

Sara Hoffmann



Directed Reading For Content Mastery Section 3 Behavior Of Gases:

Foundations of Adult Health Nursing Lois White, 2005 *Foundations of Adult Health Nursing* is written following the Nursing Process format. The clinical chapters have been revised to include Nursing Management sections for each disorder emphasizing the nurse's role in providing competent client care. Nursing Outcomes Classification (NOC) and Nursing Interventions Classification (NIC) have been identified in each Nursing Care Plan. A Case Study included in each chapter uses critical thinking questions to assist the user through the nursing process and to develop a customized plan of care for the client in the scenario. The last chapter, Integration, has been included to assist the user in understanding the complexity of human illness and how a disorder in one body system affects other body systems. Special features throughout the book are specially designed to emphasize specific information such as Life Cycle Considerations, Client Teaching, Infection Control, Safety, Community Home Health Care, Cultural Considerations, and Professional Tips. Critical Thinking Questions are included throughout the book to encourage the user to apply the chapter's content. At the end of each chapter, readers will find a Summary and Review Questions along with References, Suggested Readings, and Resources.

Bulletin of the Atomic Scientists, 1970-12 The *Bulletin of the Atomic Scientists* is the premier public resource on scientific and technological developments that impact global security. Founded by Manhattan Project Scientists, the *Bulletin's* iconic Doomsday Clock stimulates solutions for a safer world.

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Integrated Physics and Chemistry, Chapter 3 Activities Paradigm Accelerated Curriculum, 2005-01-01 Key topics: Chemical nomenclature, Lavoisier's list of elements, sulfur, diamonds, graphite, coal, medieval metals, platinum, zinc, cobalt, nickel, manganese, molybdenum, tungsten, gases in the atmosphere, air pressure and humidity, Henry Cavendish, hydrogen, nitrogen, fertilizers and explosives, dynamite, laughing gas. IPC consists of twelve chapters of text and twelve companion student activity books. This course introduces students to the people, places, and principles of physics and chemistry. It is written by internationally respected scientist-author John Hudson Tiner who applies the vignette approach which effectively draws readers into the text and holds attention. The author and editors have deliberately avoided complex mathematical equations in order to entice students into high school level science. Focus is on the people who contributed to development of the Periodic Table of the Elements. Students learn to read and apply the Table while gaining insight into basic chemistry and physics. This is one of our most popular courses among high school students, especially those who have a history of underperformance in science courses due to poor mathematical and reading comprehension skills. The course is designed for two high school transcript credits. Teachers may require students to complete all twelve chapters for two transcript credits or may select only six chapters to be completed for one transcript credit for Physical Science, Physics, or Chemistry. Compliance with state and local academic essential elements should be considered when specific chapters are selected by teachers. As applicable to local policies, transcript credit may be assigned as follows when students complete all 12 chapters: Physical Science for one credit and

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Gases Sara Hoffmann, 2013 Students will enjoy learning about solids liquids and gases while improving their reading skills Full color photographs and easy sentences introduce students to the three different types of matter This series meets both science and reading standards

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