

Chapter 6 - Lecture Worksheet 3

Using standard heats of formation, calculate the standard enthalpy change for the following reaction.



$$\begin{aligned}\Delta H^\circ_{\text{rxn}} &= \Delta H^\circ_f(\text{CH}_4, \text{g}) + 2\Delta H^\circ_f(\text{H}_2\text{O}, \text{g}) - \Delta H^\circ_f(\text{CO}_2, \text{g}) - 4\Delta H^\circ_f(\text{H}_2, \text{g}) \\ &= -75 + 2(-242) - (-394) - 4(0) \text{ kJ} = -165 \text{ kJ}\end{aligned}$$

(1) -711 kJ (2) -165 kJ (3) -15 kJ (4) 15 kJ (5) 165 kJ (6) 229 kJ (7) 711 kJ

Using average bond enthalpies estimate the standard enthalpy change for the same reaction.



$$\begin{aligned}\Delta H^\circ_{\text{rxn}} &= 2BE(\text{C}=\text{O}) + 4BE(\text{H}-\text{H}) - 4BE(\text{C}-\text{H}) - 4BE(\text{O}-\text{H}) \\ &\approx 2(728) + 4(436) - 4(413) - 4(463) \text{ kJ} = -304 \text{ kJ}\end{aligned}$$

Which is more accurate ? Why ?

ΔH° calculated from ΔH°_f is more accurate because ΔH° calculated from bond enthalpies uses AVERAGE bond enthalpies obtained from a number of different compounds in which the bonds arise and are not specific to the compound in question.

	<u>Standard Enthalpy of formation (kJ/mol)</u>
$\text{CH}_4(\text{g})$	-75
$\text{CO}_2(\text{g})$	-394
$\text{C}_2\text{H}_2(\text{g})$	+227
$\text{H}_2\text{O}(\text{g})$	-242
$\text{H}_2\text{O}(\text{l})$	-286
$\text{NH}_3(\text{g})$	-46
$\text{HCl}(\text{g})$	-92
$\text{HCl}(\text{aq})$	-167
$\text{CaCO}_3(\text{s})$	-1208
$\text{Ca}(\text{OH})_2(\text{aq})$	-1003
$\text{Ca}(\text{OH})_2(\text{s})$	-986
$\text{CaO}(\text{s})$	-635
$\text{SO}_2(\text{g})$	-297
$\text{SO}_3(\text{g})$	-396

	<u>Average Bond Enthalpies (kJ/mol)</u>
H-H	436
N-H	391
N-N	161
N=N	418
N≡N	945
O-H	463
C-O	351
C=O	728
CaO	1072
Cl-Cl	242
Cl-H	432
Br-Br	193
Br-H	366
C-H	413

Enthalpy Worksheet With Answers

Masterton



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Holt Chemistry Holt Rinehart & Winston, 2003-01-24 Educart CBSE Class 12 Chemistry One Shot Question Bank 2026
(Includes PYQs for 2025-26) , **Educart ICSE Class 10 One-shot Question Bank 2026 Chemistry (strictly for 2025-26 boards)** Sir Tarun Rupani, 2025-07-12 Fast track your Chemistry revision with this exam ready resource This One shot Question Bank by Sir Tarun Rupani is designed to help ICSE Class 10 students revise the complete Chemistry syllabus quickly and thoroughly It simplifies theory boosts numerical accuracy and ensures strong exam practice all aligned with the 2025 26 ICSE syllabus Key Features Strictly Based on ICSE 2025 26 Curriculum Complete chapter coverage including Periodic Table Chemical Bonding Acid Base Organic Chemistry and more One shot Format Each chapter includes concise concept notes chemical equations reactions and key diagrams for quick recall Complete Coverage of Question Types Includes objective short long answers equation based numerical and reasoning questions Chapterwise PYQs Included Practice with previous years ICSE board questions to understand trends and improve retention Solved Answers in ICSE Format Clear well structured solutions using proper units chemical symbols and balanced equations Smart Revision Focus Special tips to avoid common mistakes in writing reactions balancing equations and attempting numericals Why Choose This Book This Chemistry One shot by Sir Tarun Rupani is built for smart preparation whether you re revising at the last minute or practising throughout the term It helps you approach each question with clarity confidence and the precision needed to score high in the 2026 ICSE board exam *The Science Teacher* , 2008 SCC Library has 1964 cur **Mechanical Engineering News** , 1990 **Singlet Photodissociation of Ketene** Melonie Anne Teichert, 1999 **Commonly Asked Questions in Thermodynamics** Marc J. Assael, William A. Wakeham, Anthony R. H. Goodwin, Stefan Will, Michael Stamatoudis, 2011-03-14 Have you ever had a question that keeps persisting and for which you cannot find a clear answer Is the question seemingly so simple that the problem is glossed over in most resources or skipped entirely CRC Press Taylor and Francis is pleased to introduce Commonly Asked Questions in Thermodynamics the first in a new series of books that address **Enthalpy and Internal Energy**: Emmerich Wilhelm, Trevor Letcher, 2017-09-12 Containing the very latest information on all aspects of enthalpy and internal energy as related to fluids this book brings all the information into one authoritative survey in this well defined field of chemical thermodynamics Written by acknowledged experts in their respective fields each of the 26 chapters covers theory experimental methods and techniques and results for all types of liquids and vapours These properties are important in all branches of pure and applied thermodynamics and this vital source is an important contribution to the subject hopefully also providing key pointers for cross fertilization between sub areas **Commonly Asked Questions in Thermodynamics** Marc J. Assael, Geoffrey C. Maitland, Thomas Maskow, Urs von Stockar, William A. Wakeham, Stefan Will, 2022-08-05 CRC Press is pleased to introduce the new edition of Commonly Asked Questions in Thermodynamics an indispensable resource for those in modern science and engineering disciplines from molecular science engineering and

biotechnology to astrophysics Fully updated throughout this edition features two new chapters focused on energy utilization and biological systems This edition begins by setting out the fundamentals of thermodynamics including its basic laws and overarching principles It provides explanations of those principles in an organized manner using questions that arise frequently from undergraduates in the classroom as the stimulus These early chapters explore the language of thermodynamics the first and second laws statistical mechanical theory measurement of thermodynamic quantities and their relationships phase behavior in single and multicomponent systems electrochemistry and chemical and biochemical reaction equilibria The later chapters explore applications of these fundamentals to a diverse set of subjects including power generation with and without fossil fuels for transport industrial and domestic use heating decarbonization technologies energy storage refrigeration environmental pollution and biotechnology Data sources for the properties needed to complete thermodynamic evaluations of many processes are included The text is designed for readers to dip into to find an answer to a specific question where thermodynamics can provide some if not all of the answers whether in the context of an undergraduate course or not Thus its readership extends beyond conventional technical undergraduates to practicing engineers and also to the interested lay person who seeks to understand the discourse that surrounds the choice of particular technological solutions to current and future energy and material production problems Balancing Chemical Equations Worksheet

Crispin Collins,2020-09-12 Struggling with balancing chemical reaction Balancing chemical equations can look intimidating for lot of us The good news is that practice makes perfect Master balancing skill with this workbook packed with hundreds of practice problems This book is for anyone who wants to master the art of balancing chemical reactions First few chapters of this book are step by step explanation of the concepts and other chapters are for practicing problems This book help students develop fluency in balancing chemical equation which provides plenty of practice Methods to solve with the explanation Total of 550 problems to solve with answer key 450 chemical reactions to practice with answer key 100 practice problems that are needed before balancing a chemical reaction with answer key Click the Buy now button to take advantage of this book to help yourself in mastering balancing skill A Level Chemistry Multiple Choice Questions and Answers (MCQs)

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Thermodynamics Problem Solving in

Physical Chemistry Kathleen E. Murphy, 2020-03-23 Thermodynamics Problem Solving in Physical Chemistry Study Guide and Map is an innovative and unique workbook that guides physical chemistry students through the decision making process to assess a problem situation create appropriate solutions and gain confidence through practice solving physical chemistry problems The workbook includes six major sections with 20 30 solved problems in each section that span from easy single objective questions to difficult multistep analysis problems Each section of the workbook contains key points that highlight major features of the topic to remind students of what they need to apply to solve problems in the topic area Key Features Provides instructor access to a visual map depicting how all equations used in thermodynamics are connected and how they

are derived from the three major energy laws Acts as a guide in deriving the correct solution to a problem Illustrates the questions students should ask themselves about the critical features of the concepts to solve problems in physical chemistry Can be used as a stand alone product for review of Thermodynamics questions for major tests *Chemistry* Jamie King-Holden,2011 **Balancing Chemical Equations Worksheets (Over 200 Reactions to Balance)** Chris McMullen,2016-01-12 Master the art of balancing chemical reactions through examples and practice 10 examples are fully solved step by step with explanations to serve as a guide Over 200 chemical equations provide ample practice Exercises start out easy and grow progressively more challenging and involved Answers to every problem are tabulated at the back of the book A chapter of pre balancing exercises helps develop essential counting skills Opening chapter reviews pertinent concepts and ideas Not just for students Anyone who enjoys math and science puzzles can enjoy the challenge of balancing these chemical reactions Enthalpy Changes in Chemical Reactions/Hess's Law James M. Postma,Julian Roberts,J. Leland Hollenberg,2000-01-15 **Hess's Law** Lifeliqe,2019 This lesson plan covers two indirect methods of determining enthalpy change for most chemical reactions using Hess s law **Enthalpy and Internal Energy** Trevor M. Letcher,Emmerich Wilhelm,2018 **Coop Learning to Accompany Chemistry** Masterton,1997

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Table of Contents Enthalpy Worksheet With Answers

1. Understanding the eBook Enthalpy Worksheet With Answers
 - The Rise of Digital Reading Enthalpy Worksheet With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Enthalpy Worksheet With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Enthalpy Worksheet With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Enthalpy Worksheet With Answers
 - Personalized Recommendations
 - Enthalpy Worksheet With Answers User Reviews and Ratings

- Enthalpy Worksheet With Answers and Bestseller Lists
- 5. Accessing Enthalpy Worksheet With Answers Free and Paid eBooks
 - Enthalpy Worksheet With Answers Public Domain eBooks
 - Enthalpy Worksheet With Answers eBook Subscription Services
 - Enthalpy Worksheet With Answers Budget-Friendly Options
- 6. Navigating Enthalpy Worksheet With Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Enthalpy Worksheet With Answers Compatibility with Devices
 - Enthalpy Worksheet With Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Enthalpy Worksheet With Answers
 - Highlighting and Note-Taking Enthalpy Worksheet With Answers
 - Interactive Elements Enthalpy Worksheet With Answers
- 8. Staying Engaged with Enthalpy Worksheet With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Enthalpy Worksheet With Answers
- 9. Balancing eBooks and Physical Books Enthalpy Worksheet With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Enthalpy Worksheet With Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Enthalpy Worksheet With Answers
 - Setting Reading Goals Enthalpy Worksheet With Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Enthalpy Worksheet With Answers
 - Fact-Checking eBook Content of Enthalpy Worksheet With Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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