

20 • Entropy and Free Energy**PRACTICE TEST**

- Which of the following represents an increase in entropy?
 - freezing of water
 - boiling of water
 - crystallization of salt from a supersaturated solution
 - the reaction $2 \text{NO(g)} \rightarrow \text{N}_2\text{O}_2\text{(g)}$
 - the reaction $2 \text{H}_2\text{(g)} + \text{O}_2\text{(g)} \rightarrow 2 \text{H}_2\text{O(g)}$
- The enthalpy of vaporization of methanol (CH_3OH) is 35.3 kJ/mol at the boiling point of 64.2 °C. Calculate the entropy change for methanol going from a liquid to vapor.
 600. J/K·mol
 - 551 J/K·mol
 - 105 J/K·mol
 - 105 J/K·mol
 - 551 J/K·mol
- Calculate the standard entropy change for the following reaction, given that

$S^\circ[\text{Cu(s)}] = 33.15 \text{ J/K}\cdot\text{mol}$
$S^\circ[\text{O}_2\text{(g)}] = 205.14 \text{ J/K}\cdot\text{mol}$
$S^\circ[\text{CuO(s)}] = 42.63 \text{ J/K}\cdot\text{mol}$

$$\text{Cu(s)} + \frac{1}{2} \text{O}_2\text{(g)} \rightarrow \text{CuO(s)}$$
 - 195.66 J/K
 - 93.09 J/K
 - 45.28 J/K
 - 93.09 J/K
 - 195.66 J/K
- In which of the following reactions do you expect to have a decrease in entropy?
 - $\text{Fe(s)} \rightarrow \text{Fe(l)}$
 - $\text{Fe(s)} + \text{S(s)} \rightarrow \text{FeS(s)}$
 - $2 \text{Fe(s)} + 3/2 \text{O}_2\text{(g)} \rightarrow \text{Fe}_2\text{O}_3\text{(s)}$
 - $\text{HF(l)} \rightarrow \text{HF(g)}$
 - $2 \text{H}_2\text{O}_2\text{(l)} \rightarrow 2 \text{H}_2\text{O(l)} + \text{O}_2\text{(g)}$
- The formation $\frac{1}{2} \text{A}_2 + 2 \text{B}_2 + \text{C} \rightarrow \text{CAB}_4$ has an enthalpy of formation of -104 kJ and a change in entropy of -60.8 J/K at 30 °C. What is ΔG and spontaneity of the reaction?
 - 85.6 kJ, spontaneous
 - 18.3 kJ, not spontaneous
 - +18.3 kJ, spontaneous
 - +85.6 kJ, not spontaneous
 - 85.6 kJ, not spontaneous
- If ΔH and ΔS are both negative or positive, then ΔG has a _____ sign.
 - positive
 - negative
 - variable
 - large
 - no
- At what temperature would a given reaction become spontaneous if $\Delta H = +119 \text{ kJ}$ and $\Delta S = +263 \text{ J/K}$?
 - 452 K
 - 2210 K
 - 382 K
 - 2.21 K
 - 363 K
- The free energy change for a given reaction is -36.2 kJ. What is the equilibrium constant at 298 K?
 - 0.985
 - 2.22×10^6
 - 1.01
 - 8.32×10^{-7}
 - 3.25×10^6
- Given the following information, calculate ΔG° for the reaction below at 25°C:

$$\text{SnCl}_4\text{(l)} + 2 \text{H}_2\text{O(l)} \rightarrow \text{SnO}_2\text{(s)} + 4 \text{HCl(g)}$$

$$\Delta H^\circ = 133.0 \text{ kJ and } \Delta S^\circ = 401.5 \text{ J/K}$$
 - 252.6 kJ
 - 13.4 kJ
 - 13.4 kJ
 - 122.9 kJ
 - 252.6 kJ

Entropy Energy And Equilibrium Practice Test

Hugh Longbourne Callendar



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