

PRE-AP CHEMISTRY EQUATIONS AND CONSTANTS

ATOMIC STRUCTURE

$$E = h\nu \quad \lambda \nu = c \quad E = \frac{hc}{\lambda}$$

STOICHIOMETRY

$$M_{\text{molar}} = M_A + M_B + M_C + \dots$$

$$\rho = \frac{m}{V}$$

$$\text{molar} = \frac{\text{moles}}{\text{L}} \quad \text{or} = \frac{\text{mol}}{\text{volume}}$$

$$\frac{M_A}{M_A + M_B} = \frac{M_C}{M_C + M_D}$$

SOLUTIONS

$$\text{molarity} = \frac{\text{moles of solute}}{\text{volume of solution}}$$

$$M_1V_1 = M_2V_2$$

$$K_w = (10^{-7})(10^{-7})$$

$$\text{pH} = -\log (10^{-7})$$

$$\text{pOH} = -\log (10^{-7})$$

$$\text{pH} + \text{pOH} = 14$$

$$\text{pH} = 14 - \text{pOH} = 7$$

$$\text{pH} = \text{pOH} = 7$$

ENERGISTICS

$$E = mc^2 \quad (\text{mass energy eq. } 10^{16} \text{ J})$$

$$\Delta H = \text{moles} \times M_{\text{molar}}$$

$$\Delta H_{\text{total}} = \sum n_i \Delta H_f^{\circ}(\text{products}) - \sum n_j \Delta H_f^{\circ}(\text{reactants})$$

OTHER FORMULAE

$$E = \frac{hc}{\lambda}$$

$$\text{m} = \frac{\text{mass}}{\text{volume}} = \left(\frac{\text{density} \times \text{volume} + \text{density} \times \text{volume}}{\text{density}} \right) \text{mass}$$

$$\text{m} = \frac{\text{mass}}{\text{volume}} = \left(\frac{\text{density} \times \text{volume}}{\text{density} \times \text{volume}} \right) \text{mass}$$

EQUILIBRIUM

$$K_{\text{eq}} = \frac{[\text{products}]}{[\text{reactants}]}$$

$$K_{\text{eq}} = \frac{[\text{products}]}{[\text{reactants}]}$$

LANGMUIR AND LANGMUIR

$$\text{Langmuir's Equation} = 10^{-10} \times 10^{10} \text{ (particles per mole)}$$

$$K = \text{Langmuir's Constant} = 10^{-10} \times 10^{10} \text{ J/mol}$$

$$K = \text{Langmuir's Constant} = 10^{-10} \times 10^{10} \text{ J/mol}$$

$$K_{\text{eq}} = \text{Langmuir's Constant} = 10^{-10} \times 10^{10} \text{ (molecules)}^2$$

$$K_{\text{eq}} = 10^{-10} \times 10^{10} \text{ J/mol}$$

$$K_{\text{eq}} = 10^{-10} \text{ J/mol}$$

$$\text{Volume of ideal gas at } 10^6 = 1.1 \times 10^6 \text{ L/mol}$$

$$K_{\text{eq}} = 10^{-10} \times 10^{10} \text{ J/mol} = 10^{-10} \times 10^{10} \text{ J/mol} = 10^{-10} \times 10^{10} \text{ J/mol}$$

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Daniela Niemeyer



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1. Understanding the eBook Download File Memorial Hs Pre Ap Chem Dr Johnson
 - The Rise of Digital Reading Download File Memorial Hs Pre Ap Chem Dr Johnson
 - Advantages of eBooks Over Traditional Books
2. Identifying Download File Memorial Hs Pre Ap Chem Dr Johnson
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Download File Memorial Hs Pre Ap Chem Dr Johnson
 - User-Friendly Interface
4. Exploring eBook Recommendations from Download File Memorial Hs Pre Ap Chem Dr Johnson
 - Personalized Recommendations
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 - Highlighting and Note-Taking Download File Memorial Hs Pre Ap Chem Dr Johnson
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 - Participating in Virtual Book Clubs
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 - Benefits of a Digital Library
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 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Setting Reading Goals Download File Memorial Hs Pre Ap Chem Dr Johnson
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Download File Memorial Hs Pre Ap Chem Dr Johnson
 - Fact-Checking eBook Content of Download File Memorial Hs Pre Ap Chem Dr Johnson
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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