1. How did Bohr expand on Rutherford's model of the atom?

The Rutherford model of the atom was incomplete because it didn't explain how the atom's negatively charged electrons were distributed in the space surrounding the positively charged nucleus.

2. Compare the energy of an electron in the ground state and an electron in the excited state.

A valence electron in the ground state has to absorb energy in order for it to move or jump from the ground state to the excited state. This means that a ground state electron is lower energy than an excited state electron.

3. When an electron falls from a higher level to a lower level, how is the energy released?

When a valence electron absorbs energy in the form of heat or light it uses that energy to jump to an excited state (outer level). An excited (valence) electron then releases that energy in the form of light (and a little heat) and falls back to its ground state. The movement of the valence electron from the ground state to the excited state is momentary.

4. Explain how the gaseous neon atoms in a neon sign emit light.

Gaseous atom (nonmetal atoms) absorb energy in the form of high voltage electrical energy. The valence electrons of the gaseous atoms use this energy to jump from the ground state to an excited state and then emit (give off) that energy in the form of light (and a little heat).

5. List the seven colors of the visible spectrum in order of increasing energy.

Red, Orange, Yellow, Green, Blue, Indigo, and Violet

6. What is the energy difference between a photon of yellow light and a photon of violet light?

A photon of yellow light is lower energy than a photon of violet energy.

- 7. Determine the type of radiation (gamma rays, infrared waves, or radio waves) that has the:
 - a. longest wavelength radio waves
 - b. highest frequency gamma rays
 - c. greatest energy _____gamma rays _____

Electron Configuration Webquest Answers

William Fay Luder

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Integrating Information Technology into the Teacher Education Curriculum Nancy Wentworth, Rodney Earle, Michael Connell, 2013-01-11 As teaching evolves teacher education must keep up This book examines systemic reforms that incorporate new technology to improve any teacher education program While there are books that address the integration of technology into teaching curricula very few address the process for teacher education faculty and the systemic reform of a teacher education program Integrating Information Technology into the Teacher Education Curriculum Process and Products of Change provides practical examples and suggestions for teacher education departments striving to integrate new technologies into their curriculum It will help in the effort to motivate faculty to make utilizing new technology a natural strategy for the teachers they are educating It describes the creation of Design Teams at Brigham Young University's McKay School of Education funded by a PT3 grant and how these teams worked to successfully reconfigure the school's teacher preparation curricula Integrating Information Technology into the Teacher Education Curriculum examines how to compose and create a curriculum design teamincluding both teacher education and content specific methods instructors training and collaboration opportunities that focus on the infusion of technology how to facilitate alignment among a university cooperating school districts the State Office of Education and other available teacher preparation programs specific case examples of the redevelopment of teacher education courses by the instructors who teach them the process of changing a technology course required by the teacher education program the process of extending grant activities to the university s partner school districts and the State Office of Education From the editors Preparing tomorrow s teachers to use technology in schools is a complex endeavor requiring the infusion of technology into curriculum and instructional practices at all levels of the pre service program In many early teacher education programs prospective teachers took a computer literacy class separate from content methods classes and rarely engaged in real collaboration on how schoolteachers could integrate technology into authentic learning experiences By focusing merely on how to use computers technology training failed by not addressing how to teach students more effectively using a variety of technological tools What teachers need to know most is how to teach content more effectively Technology integration should cause teachers to develop different perspectives through rethinking teaching and learning Teaching with technology causes teachers to confront their established beliefs about instruction and their traditional roles as classroom teachers Electron Configuration 148 Success Secrets - 148 Most Asked Questions on Electron Configuration - What You Need to Know Dale Rich, 2014-10-15 It s a brand new Electron configuration world There has never been a Electron configuration Guide like this It contains 148 answers much more than you can imagine comprehensive answers and extensive details and references with insights that have never before been offered in print Get the information you need fast This all embracing guide offers a thorough view of key knowledge and detailed insight This Guide introduces what you want to know about Electron configuration A quick look inside of some of the subjects covered Extended periodic table 8s elements List of data references for chemical elements References for chemical elements Ferrocene Structure and bonding Coordination complex Geometry History of electromagnetic theory Lorentz and Poincare Ion Formation of polyatomic and molecular ions Valence electrons Copernicium Physical and atomic Cadmium Mol Phys Term symbol Valence electron Electron configuration Pnictogen Chemical Electronic configuration Electron configuration in molecules Group 11 elements Hund s rules Octet rule Explanation in quantum theory Group 11 element Characteristics Hassium Physical and atomic Alkali metal Extended periodic table Superactinides Neutron absorption Neutron absorbers Flerovium Nuclear stability and isotopes Lewis acid Comparison with Bronsted Lowry Theory Ununoctium Calculated atomic and physical properties Tunnel diode F block Isotopes of strontium Ytterbium Aufbau principle Atoms Energy levels Pauli exclusion principle Atoms and the Pauli principle Electron shell List of elements with electrons per shell Periodic table Layout Rutherfordium Chemical Antimony Group 8 element Lanthanum Chemical properties Extended periodic table Eka superactinides Manganese Promethium Physical properties and much more Electron Configuration as the Basis of the Periodic Table William Fay Luder, 1943

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