

# ORGANIC CHEMISTRY

## FUNDAMENTALS

### Types of Organic Compounds

**Organic chemistry** is the study of natural and synthetic materials that have carbon atoms as the key chemical feature. There are more than one million known organic compounds.

#### HYDROCARBON

**ALKANE**  
 $\text{C}-\text{H}$   
 $\text{C}-\text{C}$   
 • Ethane:  $\text{C}_2\text{H}_6$   
 • Methyl (Me):  $-\text{CH}_3$   
 • Ethyl (E):  $-\text{C}_2\text{H}_5$

#### ALKENE

• Ethene:  $\text{C}_2\text{H}_4$   
 • Diene: Two  $\text{C}=\text{C}$   
 • Triene: Three  $\text{C}=\text{C}$

#### ALKYNE

• Ethyne:  $\text{C}_2\text{H}_2$

#### AROMATIC

• Benzene:  $\text{C}_6\text{H}_6$   
 • Aro:  $\text{C}_6\text{H}_5$  (Ar)

#### -O- ADDED

##### ALCOHOL

• Methanol:  $\text{Me}-\text{OH}$  (methyl alcohol)  
 • Phenol:  $\text{Ar}-\text{OH}$   
 • Diol/Glycol: (2  $-\text{OH}$ )  
 • Glycol: (3  $-\text{OH}$ )

##### ETHER

• Ethoxyethane:  $\text{Et}-\text{O}-\text{Et}$  (diethyl ether)

##### EPOXY

• Oxirane:  $\text{C}_2\text{H}_4\text{O}$  (ethylene oxide)

##### PEROXIDE

• Acetone Peroxide:  $\text{C}_4\text{H}_8\text{O}_5$

• Benzoyl Peroxide:  $\text{C}_{14}\text{H}_{18}\text{O}_5$

#### >C=O ADDED

##### ALDEHYDE

• Methanal:  $\text{H}_2\text{CO}$  (formaldehyde)  
 • Benzaldehyde:  $\text{Ar}-\text{CHO}$

##### KETONE

• 2-propanone:  $\text{Me}-\text{CO}-\text{Me}$  (dimethyl ketone, acetone)  
 • Diketone:  $\text{R}-\text{CO}-\text{R}'-\text{CO}-\text{R}''$

#### >COO ADDED

##### CARBOXYLIC ACID

• Ethanoic Acid:  $\text{Me}-\text{COOH}$  (acetic acid)  
 • Acetate Ion:  $\text{Me}-\text{COO}^-$   
 • Benzoic Acid:  $\text{Ar}-\text{COOH}$

##### Dicarboxylic Acid

$\text{HOOC}-\text{R}-\text{COOH}$

#### ESTER

• Ethyl Acetate:  $\text{Me}-\text{CO}-\text{OEt}$   
 Other derivatives:  
 • Peroxyacid:  $\text{R}-\text{CO}-\text{OOH}$   
 • Acid Anhydride:  $\text{RCO}-\text{O}-\text{CO}-\text{R}'$

#### NITROGEN ADDED

##### AMINE

• Methyl Amine:  $\text{H}_3\text{C}-\text{NH}_2$   
 • Phenyl amine:  $\text{Ar}-\text{NH}_2$  (aniline)  
 •  $\text{R}-\text{NH}_2$  (1°),  $\text{RR}'\text{NH}$  (2°),  $\text{RR}'\text{R}''\text{N}$  (3°)

##### NITRO

$\text{R}-\text{NO}_2$

##### DIAZO

$\text{R}-\text{N}=\text{N}$

##### NITRILE

$\text{R}-\text{C}\equiv\text{N}$

• Methane Nitrile:  $\text{Me}-\text{CN}$

#### AMIDE

• Acetamide:  $\text{Me}-\text{CO}-\text{NH}_2$

#### SULFUR ADDED

• Thiol:  $\text{R}-\text{SH}$   
 • Thioether:  $\text{R}-\text{S}-\text{R}'$   
 • Disulfide:  $\text{R}-\text{S}-\text{S}-\text{R}'$   
 • Thio Ester:  $\text{R}-\text{CO}-\text{SR}'$   
 • Sulfide:  $\text{R}-\text{SO}-\text{R}'$   
 • Sulfone:  $\text{R}-\text{SO}_2-\text{R}'$   
 • Sulfonic Acid:  $\text{R}-\text{SO}_3\text{H}$

#### HALOGEN ADDED

• Haloalkane:  $\text{Me}-\text{Cl}$  chloromethane  
 • Halohydrin:  $\text{Ar}-\text{X}$   
 • Chloroalkene:  $\text{Ar}-\text{Cl}$   
 • Acyl Halide:  $\text{R}-\text{CO}-\text{X}$   
 • Aryl Halide:  $\text{Ar}-\text{X}$

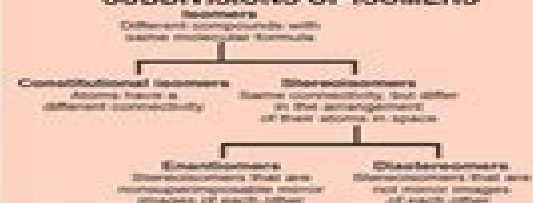
### Formulas & Isomers

- Molecular Formula:** Elemental symbols with subscripts denote the composition of a compound
- Empirical Formula:** Subscripts denote the relative elemental composition
- Graphical depiction**
  - Dash Formula:** Diagram all atoms, show bonds as dashes
  - Bond-Line Formula:** Hide H; carbon atoms are depicted as lines, but other atoms are shown explicitly



- Newman Projection:** 2-d depiction
- 3-dimensional:** Wedges of same bond denote structure

#### SUBDIVISIONS OF ISOMERS



- Constitutional Isomers:** Different bonding connectivity (e.g., rings, bonds, branching, substituent positions)
- Tautomers:** Easily interconverted structural isomers (e.g., keto-enol for ketones)
- Chiral:** Not identical with its mirror image
- Achiral:** Has a plane of symmetry (i.e., superimposable on its mirror image)
- Epimers:** Pair of diastereomers which differ only in the configuration of one atom
- More than 1 chiral center
  - n chiral centers,  $\leq 2^n$  stereoisomers
  - Meso: 2 chiral centers, 4 isomers: 3 stereoisomers, 1 achiral (mirror-plane)

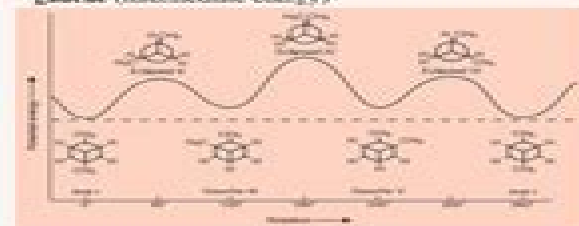


- R/S Notation:** The 4 different atoms or groups attached to a central atom are ranked (a), (b), (c), and (d), by molar mass. The lowest (d) is directed away from the viewer and the sequence of a-b-c produces clockwise (R) or counter-clockwise (S) configuration
- Chiral (Optically Active):** + or - the rotation of plane polarized light, R/S: opposite effects

- Racemic:** 50/50 mixture of stereoisomers (no net optical activity)
- Nomenclature:** Note R/S and +/- in the compound name (e.g., R (+) bromochloromethanol)
- Fisher-Projection:** Diagram depicts chiral/3-D structure
- Molecular Conformations:** Molecule exhibits structural variation due to free rotation about C-C single bond



- Newman Diagram:** Depict rotation about a C-C bond; eclipsed (high energy), anti (low energy), and gauche (intermediate energy)



### Common Terms

**Aliphatic:** Non-aromatic  
**Aromatic:** Benzene ring  
**Conjugation:** Sequence of alternating double (or triple) and single bonds  
**Dielectric Effect:** Polar solvent stabilizes ion formation  
**Exergonic Reaction:** Chemical reaction

that loses energy during the process of the reaction  
**Heterocyclic:** Non-carbon atom in the ring structure  
**Hydrocarbon:** Compound of H and C  
**Ketone:** Molecule made up of a monosaccharide bonded to a ketone

**Monosaccharide:** Carbohydrate that cannot be reduced by hydrolysis into another simple sugar  
**Olefin:** Alkene  
**Paraffin:** Alkane  
**Radioisotope:** Isotope that is radioactive

**Saturated:** Maximum # of Hs (all C-C single bonds)  
**Tetrose:** Monosaccharide with four carbon atoms  
**Unsaturated:** At least one C-C multiple bond  
**Zymology:** The study of the process of fermentation

# Fundamentals Of Organic Chemistry Study Guide

**David Kirk**



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## **Table of Contents Fundamentals Of Organic Chemistry Study Guide**

1. Understanding the eBook Fundamentals Of Organic Chemistry Study Guide
  - The Rise of Digital Reading Fundamentals Of Organic Chemistry Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Fundamentals Of Organic Chemistry Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Fundamentals Of Organic Chemistry Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Fundamentals Of Organic Chemistry Study Guide

- Personalized Recommendations
  - Fundamentals Of Organic Chemistry Study Guide User Reviews and Ratings
  - Fundamentals Of Organic Chemistry Study Guide and Bestseller Lists
5. Accessing Fundamentals Of Organic Chemistry Study Guide Free and Paid eBooks
    - Fundamentals Of Organic Chemistry Study Guide Public Domain eBooks
    - Fundamentals Of Organic Chemistry Study Guide eBook Subscription Services
    - Fundamentals Of Organic Chemistry Study Guide Budget-Friendly Options
  6. Navigating Fundamentals Of Organic Chemistry Study Guide eBook Formats
    - ePub, PDF, MOBI, and More
    - Fundamentals Of Organic Chemistry Study Guide Compatibility with Devices
    - Fundamentals Of Organic Chemistry Study Guide Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Fundamentals Of Organic Chemistry Study Guide
    - Highlighting and Note-Taking Fundamentals Of Organic Chemistry Study Guide
    - Interactive Elements Fundamentals Of Organic Chemistry Study Guide
  8. Staying Engaged with Fundamentals Of Organic Chemistry Study Guide
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Fundamentals Of Organic Chemistry Study Guide
  9. Balancing eBooks and Physical Books Fundamentals Of Organic Chemistry Study Guide
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Fundamentals Of Organic Chemistry Study Guide
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Fundamentals Of Organic Chemistry Study Guide
    - Setting Reading Goals Fundamentals Of Organic Chemistry Study Guide
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Fundamentals Of Organic Chemistry Study Guide

- Fact-Checking eBook Content of Fundamentals Of Organic Chemistry Study Guide
- Distinguishing Credible Sources

### 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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